EXTENDED ESSAY IN CHEMISTRY HL

Release of quercetin from Fe3O4 mesoporous magnetite nanoparticles (MMNPs) using external magnetic fields

# RESEARCH QUESTION

**How does the increase of the frequency of applied external magnetic fields increase quercetin release from the mesoporous magnetite nanoparticles, when using magnetic field system containing permanent magnet and solenoid with perm alloy core, connected to signal generator alternating current of 100 mA?**

# WORD COUNT: 3998

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# INTRODUCTION

Flavonoids are a broad class of polyphenolic biomolecules, with numerous hydroxyl groups, found in a variety of fruits and vegetables. They exert various biological activities including anticarcinogenic1, anti-•inflammatory2 and, antibacterial activity3. Presumably they possess remarkable therapeutic potential in preventing the onset and progression of Alzheimer's disease and in promoting cognitive performance4.

The use of flavonoids has been limited due to their poor water solubility (high hydrophobicity) and chemical instability under physiological conditions(temperature,light,pH)5. Within this work, I will investigate “**How does the increase of the frequency of applied external magnetic fields increase quercetin release from the mesoporous magnetite nanoparticles, when using magnetic field system containing permanent magnet and solenoid with perm alloy core, connected to signal generator alternating current of 100 mA?**“ Quercetin(Figure.1), a flavonoid from the

1 I. L. Martins, C. Charneira, V. Gandin, J. L. Ferreira da Silva, G. C. Justino, J. P. Telo, A. J. Vieira, C. Marzano,

A. M. Antunes, Selenium-•containing chrysin and quercetin derivatives: attractive scaffolds for cancer therapy, Medic. Chem. 58 (2015) 4250–4265.

2 M. M. Li, X. Q. Su, J. Sun, Y. F. Gu, Z. Huang, K. W. Zeng, Anti-•inflammatory ursane-• and oleanane-•type triterpenoids from Vitex negundo var. cannabifolia. J.Natural Produc. 77 (2014) 2248–2254.

3 R. Hendra, S. Ahmad, A. Sukari, M. Y. Shukor, E. Oskoueian, Flavonoid Analyses and Antimicrobial Activity of Various Parts of Phaleria macrocarpa (Scheff.) Boerl Fruit, Int. J. Mol. Sci. 12 (2011) 3422-•3431.

4 F. I. Baptista, A. G. Henriques, A. M. Silva, J. Wiltfang, O. A. da Cruz e Silva, Flavonoids as therapeutic compounds targeting key proteins involved in Alzheimer's disease, ACS Chem. Neurosci. 5 (2014) 83-•92.

5 H. Pool, D. Quintanar, J. D. Figueroa, C. M. Mano, J. E. H. Bechara, L. A. Godınez, S. Mendoza, Antioxidant Effects of Quercetin and Catechin Encapsulated into PLGA Nanoparticles, J. Nanomater. 2012 (2012) 1-•12.

subgroup of flavonols6 was investigated. Literature shows that upon therapeutic intervention with quercetin, significant neuroprotection and neuronal recovery can be achieved7.

A promising way to deliver poorly soluble drugs or other bioactive molecules is their incorporation within nanoparticles(NPs)8. Among different NPs, biodegradable NPs are gaining increased attention for their ability to serve as viable nanocarriers for site specific delivery of drugs within the body. They offer enhanced biocompatibility and convenient release profiles for a number of drugs, vaccines and biomolecules9.

However, lately biodegradable mesoporous NPs emerge as the ones having ideal properties for designed delivery system: effectively controlled particle size and surface chemistry;; enhanced permeation,flexibility,solubility and release of therapeutically active agents10. Due to their large active surface area and high pore volume, they can host diverse molecules and as result have the highest drug-•loading efficiency11. They can provide excellent physico-•chemical protection from their degradation in physiological conditions, like during endogenic enzymatic activities12.

6 M. Jazvinšćak Jembrek, L. Vuković, J. Puhović, J. Erhardt, N. Oršolić, Neuroprotective effect of quercetin against hydrogen peroxide-•induced oxidative injury in P19 neurons J. Mol. Neurosci. 47 (2012) 286-•299.

7 ibid.

8 A. Baezaab, M. Colillaab,M. Vallet-•Regí, Advances in mesoporous silica nanoparticles for targeted stimuli-•responsive drug delivery, Expert Opin. Drug Deliver. 12 (2015) 319-•337.

9 A. Mahapatro, D. K. Singh, Biodegradable nanoparticles are excellent vehicle for site directed in-•vivo delivery of drugs and vaccines, J. Nanobiotechn. 55 (2011) 1-•11.

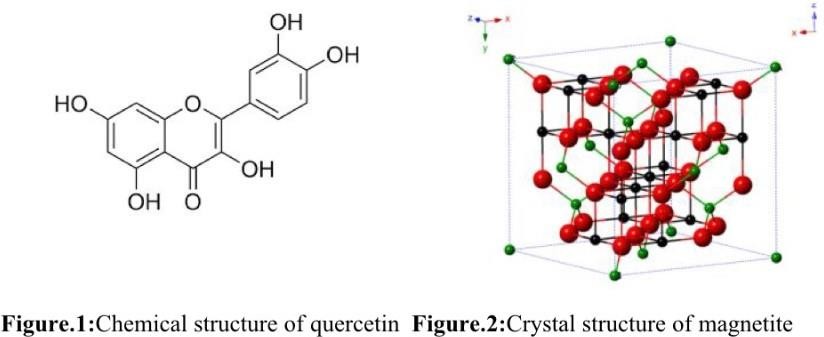
10 D. Bennet, S. Kim, Nanotechnology and Nanomaterials in: Ali Demir Sezer "Application of Nanotechnology in Drug Delivery", 2014.

11 N. Han, Y. Wang, J. Bai, J. Liu, Y. Wang, Y. Gao, T. Jiang, W. Kang, S. Wang, Facile synthesis of the lipid bilayer coated mesoporous silica nanocomposites and their application in drug delivery, Micropor. Mesopor. Mater. 219 (2016) 209-•218.

12 ibid.

As a control of their beneficial and superior flavonoid loading efficiency, Fe3O4 magnetic mesoporous nanoparticles(MMNPs) were used as responsive drug carriers to specific stimuli, such as magnetic fields to release and/or deliver drugs in a therapeutically desirable manner13. Iron oxides are widespread in nature and can be synthesized in the laboratory.

Eight iron oxides are known14, and among these, magnetite(Fe3O4) is a promising candidate. Fe3O4(Figure.2) has the face centered cubic spinel structure, based on 32O2 − ions. Fe3O4 differs from most other iron oxides in composition containing both divalent and trivalent iron. Fe3O4 has a cubic inverse structure that consists of a cubic close packed array of oxide ions, where all of the Fe2+ ions occupy half of the octahedral sites while the Fe3+ are split across the octahedral and the tetrahedral sites. In stoichiometric magnetite, the ratio is equal to Fe2+/Fe3+=0.5.



13 X. Z. Zhang, D. Q. Wu, C. C. Chu, Synthesis, characterization and controlled drug release of thermosensitive IPN-•PNIPAAm hydrogels Biomaterials 25 (2004) 3793-•3785.

14 R. M. Cornell and U. Schwertmann, The Iron Oixdes: Structures, Properties, Reactions, Occurences and Used, Weinheim: Wiley, 1996, New York.

Fe3O4 exhibits ferromagnetism at room temperature, with the saturation magnetization reaching to 92emug−1.15MMNPs become superparamagnetic at room temperature when the size of crystallites within NPs is below 15nm. However, a common phenomenon observed among superparamagnetic NPs is aggregation. Therefore, it is crucial to develop a proper protection strategy to chemically stabilize bare MMNPs against aggregation during their application. For biomedical applications, it is necessary to synthesize MMNPs that are dispersible in most biological media. Functionalization by polymer provides high colloid stability, and plays a signiﬁcant role in their bio-•distribution16. A large number of natural and synthetic biodegradable polymers is used as functionalized material of MMNPs, but here polyethylene glycol(PEG)400017 was used.

The most decisive magnetic properties are responsiveness to outer magnetic ﬁeld and magnetization, which is usually obtained from the measured hysteresis loops18(M–H) and zero-•ﬁeld cooled/ﬁeld cooled(ZFC/FC, M–T) curves. The saturation magnetization(Ms), can be determined from the hysteresis loops. In the case of superparamagnetic MMNPs, the M–H curve does not show hysteresis, and the forward and backward magnetization curves overlap.

19

15 M. Yamaura, R. L. Camilo, L. C. Sampaio, M. A. Macedo, M. Nakamura, H. E. Toma, Preparation and characterization of (3-•aminopropyl) triethoxysilane-•coated

magnetite nanoparticles J. Magn. Magn. Mater. 279 (2004) 210.

16 X. Q. Yang, J.J. Grailer, I. J. Rowland, A. Javadi, S. A. Hurley, D. A. Steeber, S. Q. Gong, Multifunctional SPIO/DOX-•loaded wormlike polymer vesicles for cancer therapy and MR imaging Biomater. 31 (2010) 9065.

17 A. J. Cole, A. E. David, J. Wang, C. J. Galbán, V. C. Yang, Magnetic brain tumor targeting and biodistribution of long-•circulating PEG-•modiﬁed, cross-•linked starch-•coated iron oxide nanoparticles Biomater. 32 (2011) 6291.

18 hysteresis loop-• a four quadrant B-•H graph where hysteresis loss, coercive force and retentively of s magnetic material are obtained -• “Hysteresis Loop”. 24th February 2012. https:/[/www.electrical4u.com/hysteresis-•l](http://www.electrical4u.com/hysteresis-)oop/. 20th June 2018.

19 W. Wu, X. H. Xiao, S. F. Zhang, H. Li, X. D. Zhou, C. Z. Jiang, One-•pot reaction and subsequent annealing to synthesis hollow spherical magnetite and maghemite Nanocages, Nanoscale Res. Lett. 4 (2009) 926.

An external magnetic ﬁeld causes reorientation of moments of individual particle along the applied ﬁeld at low temperatures. By applying an external alternating magnetic ﬁeld, reorientation of the nanoparticles and increase of precession of the nanoparticles appear causing increase in drug release from nanoparticles20. The dependent variable is amount of quercetin released, the independent frequency of applied external magnetic fields and the controlled ones are time,temperature,pressure,magnet type,solenoid type,permanent magnetic field,current of 100mA, measuring devices, apparatus, reactants(concentration,volume and mass of reactants:PEG4000,PBS,EtOH,quercetin,NH4Ac,FeCl3×6H2O,ethylene glycol).

# HYPOTHESIS:

**Applied external magnetic fields(permanent&alternating) will increase quercetin release from MMNPs depending on their response to an external magnetic ﬁeld**.

If a magnetic material is placed in a magnetic ﬁeld of strength *H*, the individual atomic moments in the material contribute to its overall response, the magnetic induction21:

*B* = 0(*H* + *M* )

where *µ*0 is permeability of free space, and the magnetization *M*

*M* = *m*

*V*

20 W. Wu, X. H. Xiao, F. Ren, S.F. Zhang, C. Z. Jiang, A comparative study of the magnetic behaviour of single and tubular clustered magnetite nanoparticles J. Low Temp. Phys. 168 (2012) 306.

21 Q A Pankhurst1, J Connolly2, S K Jones3 and J Dobson, Applications of magnetic nanoparticles in biomedicine, J. Phys. D: Appl. Phys. 36 (2003) R167–R181.

is the magnetic moment per unit volume, where *m* is the magnetic moment on a volume *V* of the material. All materials are magnetic to some extent, but their response depends on their atomic structure and temperature. The magnetization of materials depends on their volumetric magnetic susceptibility(*χ*) where

*M* = *H*

describes the magnetization induced in material by magnetic ﬁeld of strength *H*. Upon application of a magnetic ﬁeld, diamagnetic and paramagnetic materials develop negative and positive magnetization, respectively, which disappears upon removal of the ﬁeld. The susceptibility gives rise to the sigmoidal shape of the *M–H* curve, with *M* approaching a saturation of magnetization at large values of *H.* In ferromagnetic and ferrimagnetic materials, a hysteresis appears showing an irreversibility in the magnetization process, related to intrinsic effects. The anhysteretic but still sigmoidal *M–H* curve is observed by superparamagnetism. Therefore, the measurement of the *M–H* curve can prove the superparamagnetism of the materials.

When a superparamagnetic is removed from a magnetic ﬁeld, its magnetization decreases back to zero due to the ambient thermal energy of its environment. This process corresponds either to the physical rotation of nanoparticles themselves(“Brownian rotation”), or rotation of atomic magnetic moments within each nanoparticle(“Neel relaxation”). The Brownian rotation causes the reorientation of nanoparticles which caused increase in drug release from nanoparticles22.

In relation to the above-•described theory of magnetization, I synthesized superparamagnetic mesoporous magnetite nanoparticles filled with quercetin. If the synthesized nanoparticles are superparamagnetic(which will be checked by measuring the M-•H curve), the application of

22 W. Wu, X. H. Xiao, F. Ren, S.F. Zhang, C. Z. Jiang, A comparative study of the magnetic behaviour of single and tubular clustered magnetite nanoparticles J. Low Temp. Phys. 168 (2012) 306.

an external alternating magnetic field will cause rotation of the nanoparticles of the same frequency as the applied external magnetic field. As the frequency of the alternating field increases, the rotation of nanoparticles will increase. The rotation of nanoparticles should be sufficiently high to overcome the adhesion and/or electrostatic forces by which the quercetin is bound to the nanoparticle surface. This will result in increased quercetin release compared to its release without the applied external alternating magneticfield.

**METHODS**23

# Synthesis and characterization of magnetic mesoporous nanoparticles(MMNPs)

# Synthesis of MMNPs

Modified solvothermal reaction was used for preparation of mesoporous magnetic nanoparticles24,25. 1.35g FeCl ×6H O, 3.85g of NH Ac, and 0.50g of PEG4000 Da were dissolved

3 2 4

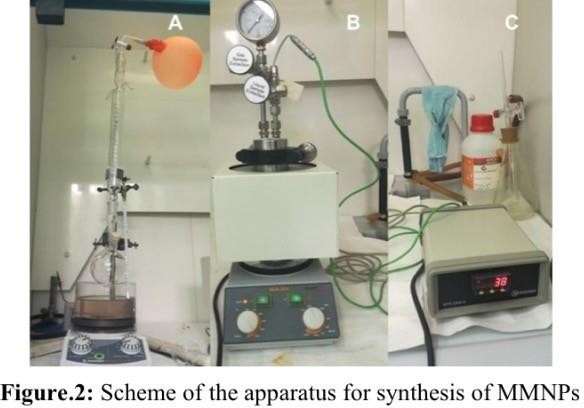
in 70mL of ethylene glycol. The mixture was placed in round two-•necked flask, volume 250mL equipped with Graham condenser and guard tube(Figure.2-•partA). Whole apparatus was placed on a hotplate-•stirrer equipped with silicone oil bath. The mixture was stirred vigorously for 1h at 160°C under the protection of N2 to form a homogeneous brownish solution and then transferred into a Teflon-•lined stainless-•steel autoclave(Figure.2-•partB) which was heated to 200°C and maintained for 16.5h. Then it was cooled to room temperature(Figure.2-•partC). The black MMNPs were then rinsed three times

23 For all materials and apparatus used observe the list of materials and apparatus in Appendix 1.

24 C: C. Wang, C. Multi-•Functional Thermosensitive Composite Microspheres with High Magnetic Susceptibility Based on Magnetite Colloidal Nanoparticle Clusters, Langmuir 26 (2010) 1674–1679.

25 B. Luo, S. Xu, W. F. Ma, W. R. Wang, S. L. Wang, J. Guo, W. L. Yang, J. H. Hu, C. C. Wang, Fabrication of Magnetite Hollow Porous Nanocrystal Shells as Drug Carrier for Paclitaxel. J. Mater. Chem. 20 (2010) 7107–7113.

with ethanol to remove solvent. The MMNPs were separated from the supernatant by using centrifuge during each rinsing step.



# Characterization of MMNPs

# Field emission scanning electron microscope(FE-•SEM)

Field emission scanning electron microscope JSM-•7000F was used for observation of particle morphology. The FE-•SEM was connected to the EDS/INCA350 energy dispersive X-•ray analyser for elemental analysis. Samples dispersed at an appropriate concentration were cast onto a glass sheet at room temperature and imaged.

# The MMNPs size distribution

The size distribution was determined using Image-•J by measuring diameters of 500 NPs based on FE-•SEM images and presented as a histogram of the nanoparticle diameters.

# X-•ray powder diffraction X-•ray diffraction(XRPD)

XRPD is very important in research of new materials. Here, X-•ray diffraction in polycrystalline is used to determine crystalline size, crystal and amorphous material differentiation and to solve and crystallize the crystalline structure. In diffraction structural analysis, monochromatic x-•ray radiation with small wavelengths is used in range from λ=0.05 to 0.25nm. Since the λ of X-•rays approximates the size of the atoms, this radiation is suitable for determining the structural arrangement of atoms and molecules of different materials. The position of the diffraction maximum is determined by crystal grating, size and form of the unit grid, and intensity of the diffraction peak atoms of the atom and their spatial deployment in the unit cell according to the requirements of the symmetry, i.e. the crystal structure. In addition to determining the position of diffraction lines that are directly related to size and shape of the unit grid, a lot of additional information is obtained from the data which affects intensity of the individual maximum lines. The structural features of prepared sample were studied and characterized by powdered X-•ray diffraction at room temperature using a Philips MPD 1880 diffractometer with monochromatic CuK a radiation (λ=0.1541nm). All samples were recorded at angle 2-•15 in range of 10-•70° with a 0.02° step with a fixed time of 10s per step.

# Brunauer-•Emmett-•Teller(BET) analysis for MMNPs porosity determination

Nitrogen adsorption-•desorption measurements were performed on an ASAP2020 accelerated surface area analyser at 77K. Before measuring, samples were degassed in a vacuum at 120°C for at least 6h.

# Magnetic characterization

To confirm the superparamagnetic property of synthesized MMNPs, magnetization measurements were performed. Magnetization of powder samples of MMNPs was measured with a MPMS-•5 commercial magnetometer equipped with a superconducting quantum interferometer device(SQUID). The measurements of samples were corrected by taking into account the ampoule and temperature independent contributions. In addition, the field dependence of the magnetization(M(H)), including magnetic hysteresis loops, was measured at temperature 290K in fields up to 10kOe.

# Fourier-•transform infrared spectroscopy(FTIR spectroscopy)

FTIR spectroscopy is a technique used to obtain an infrared spectrum of a solid, liquid or gas. An FTIR spectrometer simultaneously collects high-•spectral-•resolution data over a wide spectral range. This confers a significant advantage over a dispersive spectrometer, which measures intensity over a narrow range of wavelengths26. FTIR spectroscopy is named for Fourier transform(mathematical process), required to convert raw data into the actual spectrum. FTIR spectra were measured on an ABB Bomem MB102 spectrometer, equipped with CsI optics and a DTGS detector. All spectra were collected with a nominal resolution of 4cm-•1 and 32 scans at 25°C. The samples were dried and mixed with KBr to be compressed to a plate for measurement.

26 P. Atkins, L. Jones, L. Laverman, Chemical Principles, 6th Edition, W.H. Freeman and Company, 2013, New York, 105-•106.

# Zeta potential(*ξ*) of MMNPs

The zeta(*ξ*) potential of MMNPs was measured using a Zetasizer Nano ZS equipped with a green laser(532nm) using the M3-•PALS technique. All measurements were conducted at 25°C. Data processing was done by the Zetasizer software 6.32. Results are reported as an average value of 3 independent measurements.

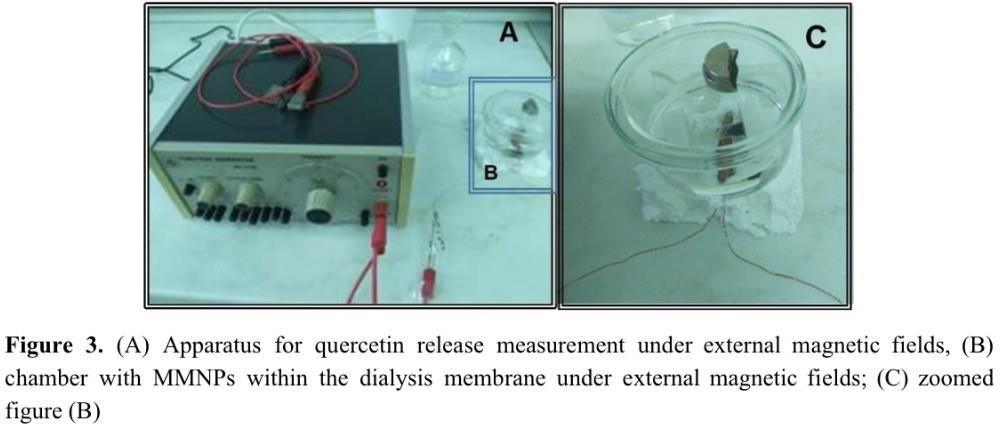
# Release of quercetin with and without external magnetic fields

The magnetic field system was set up from permanent magnet(rare earth) and solenoid with perm alloy core at distance of approximately 7cm and connected to signal generator alternating current of 100mA(Figure.3). Among the two magnetic fields, a bottle with the sample was placed. Defining the Oxy plane as the surface of the liquid, the permanent field was along the Oz axis and the alternating field along the Ox axis. In all experiments weak fields were applied: the strength of the magnetic field was 0.19T. Quercetin tests the release kinetics under permanent and alternating magnetic field. Three series of measurements of 4h duration with the field having frequencies 0.1Hz, 1Hz and 10Hz and one without the field were performed.

The release of quercetin from MMNPs(60mg) placed within Standard RC Membrane in 30mL mixture EtOH/PBS (vol.50/50) was quantified by UV absorption measurements (Bio spectrophotometer,10mm quartz cuvettes) of the supernatant solution (1.5mL) during 4h. Each aliquot of measured supernatant was replaced with same aliquot (1.5mL) of fresh mixture EtOH/PBS(vol.50/50) for maintaining the volume of the supernatant constant. Temperature in the measuring compartment was controlled and maintained at 25°C. The calibration curve was drawn by dissolving different amounts of quercetin in mixture

EtOH/PBS(vol.50/50) and after filtration of supernatant through filter (F2613-•3, PTFE 0.45µm) measuring the peak maximum in the UV absorption spectra(*λ*max= 375nm for quercetin). The linearity of calibration was found to be valid from 1×10-•6moldm-•3 to

1×10-•4moldm-•3 with correlation coefficients for quercetin all approaching to 1.00.



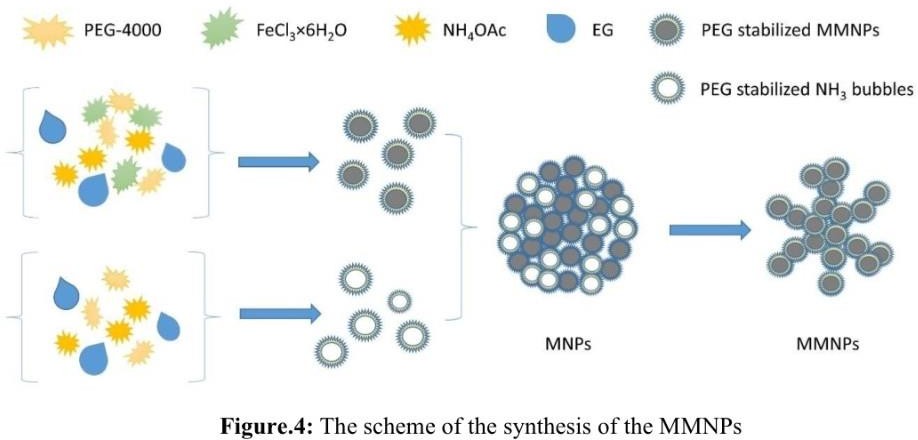
# RESULTS AND DISCUSSION

# Synthesis of MMNPs

Solvothermal reaction was performed in the formation of submicrometer MMNPs to improve their effectiveness in drug loading and decrease costs. The Wang group described the solvothermal synthesis of magnetic clusters stabilized with poly-•(acrylic acid)(PAA) which is very expensive27. Magnetic clusters with high magnetization that are able to host hydrophobic therapeutic quercetin were obtained. Therefore, this procedure was improved by replacing the PGA stabilizer with biodegradable PEG-•4000, thus improving biocompatibility,decreasing

27 B. Luo, S. Xu, A. Luo, W. R. Wang, S. L. Wang, J. Guo, Y. Lin, D. Y. Zhao, Mesoporous Biocompatible and Acid-•Degradable Magnetic Colloidal Nanocrystal Clusters with Sustainable Stability and High Hydrophobic Drug Loading Capacity, ACS Nano 5 (2011) 1428–1435.

the costs of synthesis of MMNPs,and extending their application to hydrophobic drugs. The proposed mechanism of the synthesized MMNPs is shown on Figure.4.

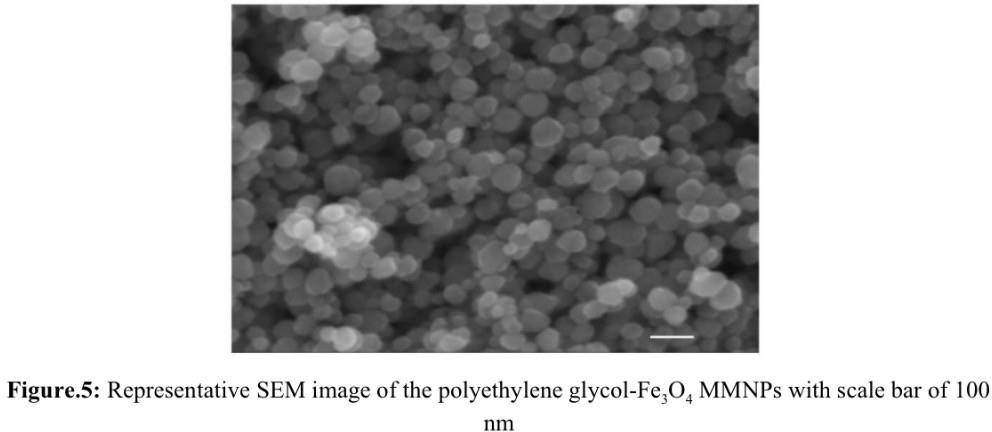


# Characterization of synthesized MMNPs

# Field emission scanning electron microscope(FE-•SEM) and size distribution

As shown in Figure.4.(scheme of synthesis process) and Figure.5(representative SEM image), the mesoporous MMNPs consisting of primary Fe3O4 nanocrystals are almost uniform both in size and shape. The size distribution of synthesized MMNPs showed their mean size of 58nm with standard deviation of 18nm(Figure.6). 28

28 For the frequency table observe Table.2 Appendix 2, and for the standard deviation, highest and lowest diameters as well as mean observe Appendix 2



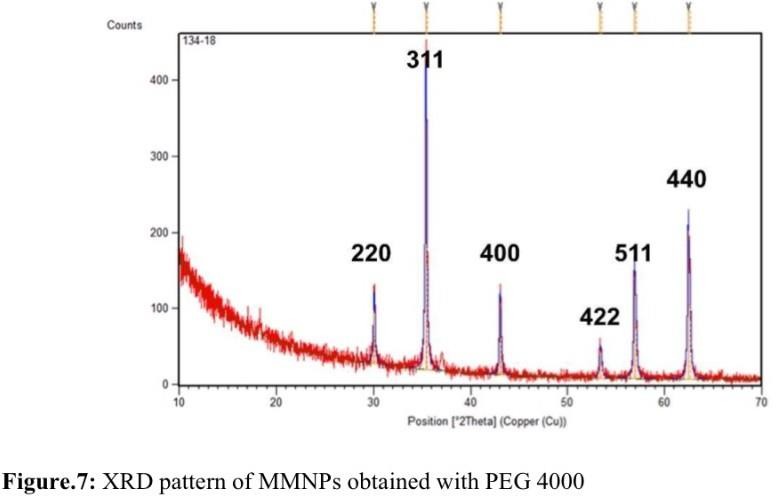


# X-•ray powder diffraction X-•ray diffraction(XRPD)

The powder X-•ray diffraction(XRD) patterns of the as-•prepared MMNPs(Figure.7) showed

that all diffraction peaks can be well indexed to the magnetic cubic structure of Fe3O4 29

exhibiting many magnetite nanocrystals whose sizes are around 35nm.



# Brunauer-•Emmett-•Teller(BET) analysis for MMNPs porosity determination

Nitrogen adsorption-•desorption isotherms of MMNPs obtained with 500mg of PEG confirmed their mesoporosity. The BET surface area, pore size and total pore volume are calculated to be 16.69m2 g-•1, 25.93nm and 0.106cm3g-•1, respectively, strongly supporting the fact that the MMNPs have mesoporous structure.30

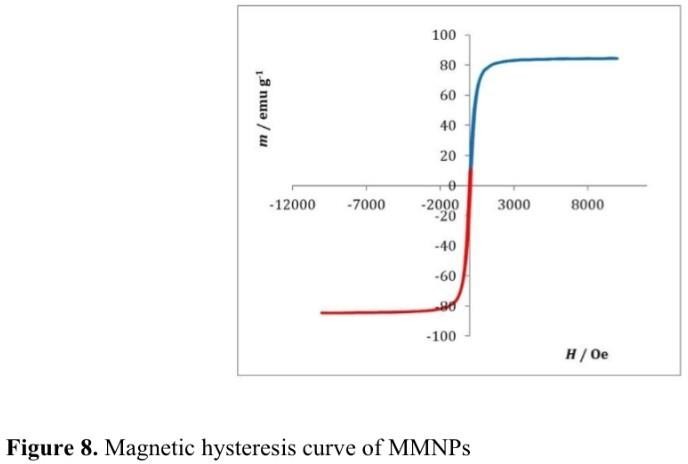
29 P. Kumar, C. Joshi, A. Barras, B. Sieber, A. Addad, L. Boussekey, S. Szunerits, R. Boukherroub, S. L. Jain, Core–shell structured reduced graphene oxide wrapped magnetically separable rGO@CuZnO@Fe3O4 microspheres as superior photocatalyst for CO2 reduction under visible light, Appl. Catal. B 205 (2017) 654-

* 665.

30 For experimental conditions performed under MMNPs observe Table.3 in Appendix 3

# Magnetic characterization

Magnetic characterization using a magnetometer at 290K indicated that the MMNPs synthesized with 500mg PEG have saturation magnetization value of 84.5emug-•1(Figure.8). Measurement of magnetisation of synthesized MMNPs confirmed their superparamagnetic properties,thus confirming their further application in release studies under applied magnetic fields.



# Loading of quercetin into MMNPs

The significant surface area, high pore volume and pore size of synthesized MMNPs ensure access to their potential to be promising drug delivery vehicles for quercetin. Loading of quercetin was performed in pure ethanol. 30mg MMNPs was added to 30mL quercetin saturated solution. The suspension of MMNPs are stirred for 24h to allow diffusion into pores. Quercetin loaded MMNPs were separated by applying centrifuge and dried in desiccator overnight. Successful loading of quercetin was conﬁrmed by UV/VIS spectroscopy, zeta potential measurements and FTIR spectroscopy.

# UV/VIS spectroscopy

The Beer-•Lambert law31is applied to determine the amount of quercetin loaded into MMNPs. UV/VIS spectroscopy was employed to directly measure quercetin concentration loss in pure EtOH supernatant above synthesized MMNPs measuring the absorbance at wavelength *λ=375*nm. The Beer-•Lambert law is the linear relationship between absorbance and concentration of an absorbing species, written as:

*A* = ɛ*bc*

where *A* is measured absorbance,ɛ( ) wavelength-•dependent molar absorptivity coefficient with units of M-•1cm-•1, b path length, and *c* molar concentration.

Compared with quercetin concentration supernatant before adding the synthesized MMNPS, the concentration loss was determined using calibration curve in pure EtOH(Figure.9). The coefficient of determination was 0.9978, and the determined molar absorption coefficient of quercetin at temperature 298K and *λ*(375nm) is 19131mol-•1Lcm-•1. The loading efficiency(LE)

32, the ability of the material to entrap a certain active substance is defined as33:

*Loading efficiency* (%) = 100×*mquercetin in MNNPs*

*m*

*synthesized MMNPs*

Using this equation, the LE was determined to be (32.47±17.28)%34

independent experiments shown in Table.1.

calculated from 8

31 “Beer-•Lambert Law”. [http://life.nthu.edu.tw/~labcjw/BioPhyChem/Spectroscopy/beerslaw.htm, 16th](http://life.nthu.edu.tw/~labcjw/BioPhyChem/Spectroscopy/beerslaw.htm%2C16th) November 2018.

32 “Encapsulation efficiency”. Research Gate. https:/[/www.researchgate.net/post/How\_do\_I\_calculate\_encapsulation\_efficiency.](http://www.researchgate.net/post/How_do_I_calculate_encapsulation_efficiency) 16th November,2018

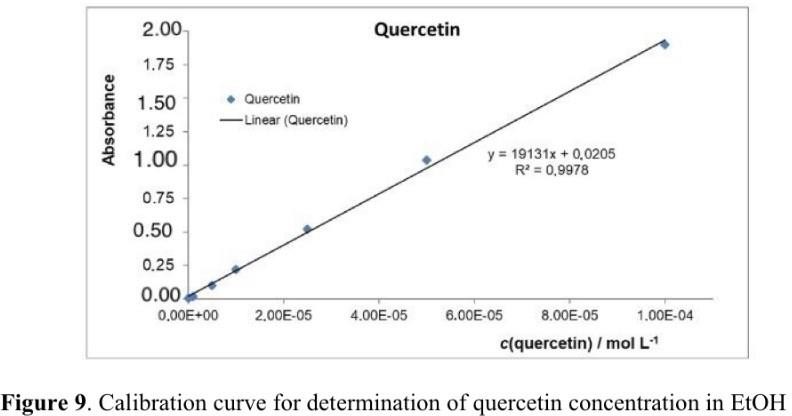
33 Notes written during the stay at the Ruder Boskovic Institute (Information given by Suzana Segota and other fellows)

34 The error is displayed as a standard deviation as to be more precise. “The advantage of using s to quote uncertainty in a result is that it has the same units as the experimental data.” “Chemistry Dictionary.” *Chemicool*, [www.chemicool.com/definition/standard\_deviation.html.](http://www.chemicool.com/definition/standard_deviation.html)

This indicates a considerable amount of quercetin loaded into MMNPs, reﬂecting their potential as drug delivery carriers.

|  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- |
| Experiment number | m (Fe3O4)  / mg ± 0.01 | c (querceti n, saturated solution)  / mol dm-•3 ± 0.00001 | V  (querceti n, saturated solution)  / mL ± 0.1 | c (querceti n, supernat ant) / mol dm-•3  ± 0.0001 | c (querceti n, loaded) / mol dm-•3  ± 0.00001 | m (kverc, loaded) / g  ± 0.0000001 | Loading efficiency  / % |
| 1 | 60.02 | 0.02407 | 31.5 | 0.0232 | 0.00087 | 0.0082827 | 13.75 |
| 2 | 60.46 | 0.02407 | 31.5 | 0.0220 | 0.00204 | 0.0194216 | 32.12 |
| 3 | 60.44 | 0.02080 | 31.5 | 0.0175 | 0.00330 | 0.0314174 | 51.98 |
| 4 | 60.24 | 0.02080 | 31.5 | 0.0179 | 0.00290 | 0.0276093 | 45.83 |
| 5 | 60.52 | 0.02140 | 31.5 | 0.0198 | 0.00160 | 0.0152327 | 25.17 |
| 6 | 60.78 | 0.01973 | 31.5 | 0.0183 | 0.00140 | 0.0133286 | 21.93 |
| 7 | 60.57 | 0.01973 | 31.5 | 0.0190 | 0.00080 | 0.0073307 | 12.10 |
| 8 | 60.14 | 0.02140 | 31.5 | 0.0178 | 0.00360 | 0.0342736 | 56.98 |

*Average loading ef f iciency* = (32.47±17.28)% Table.1: Quercetin loadings and loading efficiency



# FTIR spectroscopy

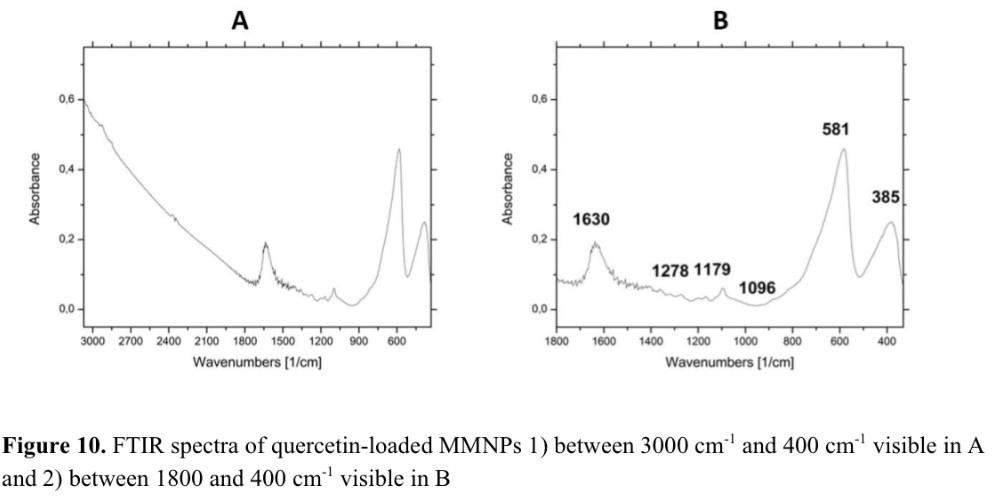
Successful loading of quercetin was conﬁrmed by FTIR spectra(Figure.10). The PEG-•modiﬁed MMNPs show peak at 1630cm-•1 that may indicate the existence of Fe–O stretching and exhibit characteristic peak of magnetite at around 581cm-•1.35Band at 1100cm−1

indicates the presence of PEG on MMNPs36. For quercetin loaded MMNPs, bands at 1278cm-•1 and 1179cm-•1 were attributable to the C–O stretching in the aryl ether ring and the C–O stretching in phenol37.

35 M. Chirita, R. Banica, A. Ieta, I. Grozescu, uperparamagnetic Unusual Behavior of Micrometric Magnetite Monodisperse Monocrystals Synthesized by Fe-•EDTA Thermal Decomposition, Particul. Sci. Technol. 30 (2012) 354–363.

36 T. P. T. Dao, T. H. Nguyen, V. V. To, T. H. Ho, T. A. Nguyen, M. C. Dang, A new formulation of curcumin using poly(lactic-•co-•glycolic acid)—polyethylene glycol diblock copolymer as carrier material, Adv. Nat. Sci. Nanosci. Nanotechnol. 5 (2014) 035013-•035020.

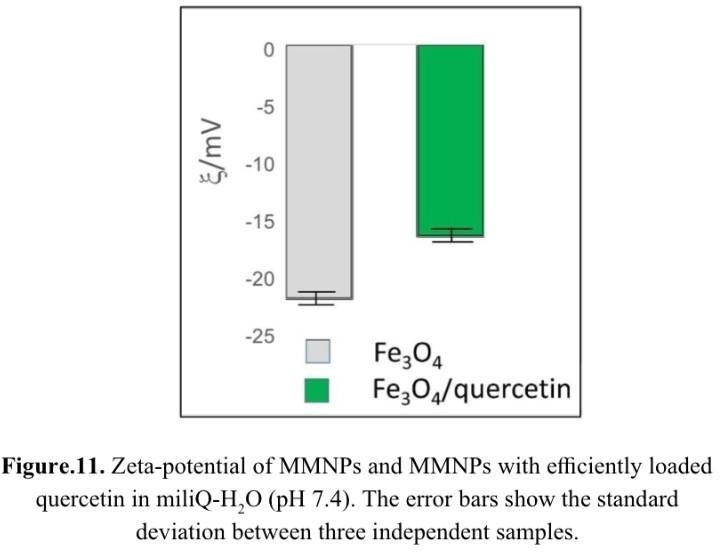
37 M. Catauro, F. Papale, F. Bollino, S. Piccolella, S. Marciano, P. Nocera, S. Paciﬁco, Silica/quercetin sol–gel hybrids as antioxidant dental implant materials, Sci. Technol. Adv. Mater. 16 (2015) 035001-•035012.



# Zeta-•potential measurements

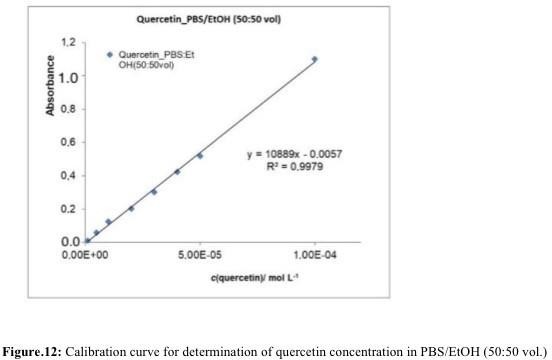
The zeta-•potential measurements for MMNPs in mili-•Q H2O at 25 °C showed that after loading quercetin, the *ξ-•*potential of pure Fe3O4 MMNPs changes from negative, *ξ*=(−22.32±0.8)mV (Fe3O4) to less negative *ξ*=(-•17±1)mV (Fe3O4/quercetin). The surface charge of the Fe3O4/quercetin is then primarily determined by the OH-•groups of quercetin

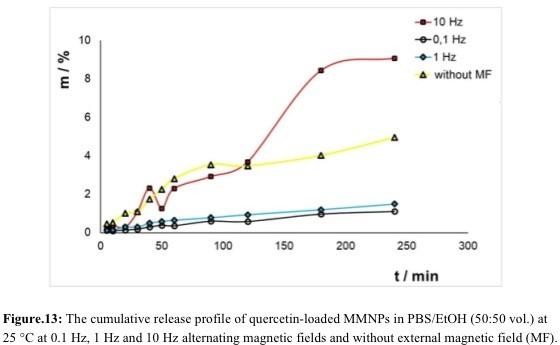
causing different charge potential relative to pure Fe3O4 MMNPs. The change in *ξ-•*potential of MMNPs before and after their exposure to feeding solution confirms successful quercetin loading(Figure.11).



# Release of quercetin with and without external magnetic fields

Quercetin tests the release kinetics without and under an external stimulus in the form of combination of permanent(0.19T) and oscillating magnetic fields (0.1Hz,1Hz and 10Hz) of 4 hour duration. To determine the amount of quercetin released from MMNPs, UV/VIS spectroscopy was employed to directly measure quercetin concentration increase in supernatant mixture of phosphate buffer solution and ethanol(PBS/EtOH), volume fraction 50:50, above quercetin-•loaded MMNPs measuring absorbance at wavelength *λ=375*nm. Compared with quercetin concentration in supernatant in time intervals of 4h, quercetin concentration increase was determined using calibration curve in PBS/EtOH (vol.50/50) (Figure.12). The coefficient of determination was 0.9979, and the determined molar absorption coefficient of quercetin at temperature 298K and *λ*=375nm is 10889mol-•1Lcm-•1. In early stage of release, by applying alternating magnetic field 0.1Hz and 1Hz, the release of quercetin is very low and practically the same(1% after 4h). In early stage of release, a burst

was observed lasting first three hours without and by applying external stimulus in first 60min. After 120min release of quercetin reached slightly over 3% for both without external magnetic fields and for combination of permanent and alternating magnetic field of 10Hz. In last, the release was increased after 3h twofold(8%) in comparison to release without applied magnetic fields(4%)(Figure.13).

38

Even very low-•frequency oscillating fields in combination with permanent magnetic field 0.19T significantly influence the release profiles. Strength of the permanent magnetic field and frequency of the alternating field is related to the behaviour of MMNPs in viscous medium. MMNPs rotate because of the torque due to the alternating magnetic field. The overall effect is weakening of interparticle interactions due to now enlarged interparticle distances and quicker and smoother release of the quercetin molecules from the MMNPs’ interior. The rotation of the nanoparticles is supreme under relative strong permanent magnetic field of 0.19T. The final release kinetics depends on the contributions of two opposite effects. By combination of permanent magnetic field 0.19T, and alternation magnetic field of 10Hz, the release is increased, while the combinations of the same permanent magnetic field with two lower frequencies,the release kinetics were decreased. Therefore, it is demonstrated that by combining both permanent and alternating magnetic fields, fine tuning of quercetin release can be adjusted.

38 For the results table look at Appendix 2.

# CONCLUSION AND EVALUATION

From the results, it was determined that the MMNPs consisting of the primary Fe3O4 nanocrystals are almost uniform in size and shape,showing size of (58±18)nm. The powder X-•ray diffraction patterns of MMNPs confirmed the magnetic cubic structure of Fe3O4 exhibiting many magnetite nanocrystals with sizes around 35nm. Nitrogen adsorption-•desorption isotherms of MMNPs confirmed mesoporosity of MMNPs revealing the 16.69m2g-•1 surface area, 25.93nm pore size and 0.106cm3g-•1 total pore volume. Measurement of magnetisation of MMNPs showed no hysteresis in M-•H curve revealing saturation magnetization value of 84.5emug-•1 confirming their superparamagnetic properties and further application in release studies under applied magnetic fields.

Loading of quercetin in MMNPs was conﬁrmed by UV/VIS spectroscopy, zeta-•potential measurements and FTIR. The loading efficiency is (32±17)% indicating that a considerable amount of quercetin was loaded into MMNPs. After loading with quercetin, zeta-•potential of pure Fe3O4 MMNPs changes from negative,*ξ*=(−22.32±0.8) mV (Fe3O4) to less negative

*ξ*=(-•17±1) mV (Fe3O4/quercetin). The quercetin release kinetics without and under an external stimulus in the form of combination of permanent (0.19T) and oscillating magnetic fields (0.1Hz, 1Hz and 10Hz) of 4 hours’ duration were examined. By applying alternating

magnetic field 0.1Hz, 1Hz and 10Hz, release of quercetin is increased, but the release is lower and practically the same (1% after 4h) at frequencies of 0.1Hz and 1Hz than release without applied external stimulus. In contrast, by applying the alternating magnetic fields of 10Hz, in early stage of release, a burst was observed lasting first 3h without and by applying external stimulus in first 60 minutes. After 120min release of quercetin reached slightly over 3% for both without external magnetic fields and for combination of permanent and

alternating magnetic field of 10Hz. In last, the release was increased after 3h twofold(8%) in comparison to release without applied magnetic fields(4%). Only in the case of quercetin release with combination of alternating magnetic field of 10Hz and permanent magnetic field of 0.19T is the hypothesis of increasing quercetin release confirmed, but not at the two other frequencies of 0.1Hz and 1.0Hz. However, this indicates high potential of mesoporous materials as drug carriers since MMNPs posses high magnetization, narrow size distribution, high loading efficiency and prominent biocompatibility. Fe3O4-•MMNPs could be considered as universal and stable drug delivery material particularly able to load and release with respectable efficiency quercetin having specific physico-•chemical and/or structural properties. The alternating magnetic field is neither the single nor decisive factor in release kinetics of quercetin.

The results of quercetin release from the MMNPs are in accordance with initial hypothesis. The deviations from expected results regarding quercetin release without applied alternating and permanent magnetic fields are not expected but the strength of the permanent magnetic field and frequency of the alternating field is certainly related to the behaviour of MMNPs in viscous medium. Diversely, the rotation of nanoparticles is supreme under relative strong permanent magnetic field of 0.19T. The final release kinetics thus depends on the contributions of two opposite effects.

The reason for the deviation from hypothesis could lay in the synthesis of the MMNPs since it resulted in insufficient mass of MMNPs for all experiments at once, it was necessary to repeat the synthesis several times.

Everything was performed in controlled conditions of temperature and pressure. Concentrations of all reactants were controlled. However, the synthesized MMNPs possessed chemical and structural properties that could vary. Besides, the performed experiment of the

release of quercetin from MMNPs took place in a chemistry laboratory that was temperature-

* conditioned. The reaction system in which MMNPs were located was not thermally insulated and stabilized. Therefore, there is a possibility of discontinuing the release medium. It is possible to repair the experimental conditions so that only the release takes place under temperature strictly controlled conditions.

An extension would include adding thermogravimetric analysis to the procedure and using other flavonoids, thus comparing the results.39

39 View Appendix 3 for detailed extension.

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40 For official letter written by external supervisor view Appendix.

# APPENDIX APPENDIX.1.

LIST OF MATERIALS AND APPARATUS

In this work many apparatuses were required for the applied procedures in the previous section.

# Synthesis of MMNPs

REAGENTS AND CHEMICALS REQUIRED:

* Ethanol, 98%, TTT d.o.o. Sveta Nedjelja, Croatia
* Iron(III) Chloride Hexahydrate, 97%, Alfa Easar, Germany
* Ammonium acetate, ≥ 99 %, Fluka, Switzerland
* PEG -• Polyethylene glycol, 4000 Da, ≥ 99 %, Sigma Aldrich, USA
* Ethylene glycol G.R., Lach-•ner, Czechia
* Silicone oil, high temperature, Acros Organics France, 500 mL
* Deionized water (Milipore Q water)

# Characterization of MMNPs

APPARATUS REQUIRED



* Analytical balance, Phoenix Instrument, Garbsen, Germany



* Philips MPD 1880 diffractometer with monochromatic CuK a radiation (λ  =  0.1541 nm)
* Field emission scanning electron microscope (FE-•SEM) JSM-•7000F (JEOL)ASAP2020 (Micromeritics, USA) accelerated surface area analyser



* MPMS-•5 commercial magnetometer equipped with a superconducting quantum interferometer device (SQUID)



* FTIR ABB Bomem MB102 spectrometer, equipped with CsI optics and a DTGS detectorVortex V1 S000, IKA, Staufen, Germany
* ASAP 2020, Micromeritcs, Germany REAGENTS AND CHEMICALS REQUIRED
* Potassium bromide, per IR spectroscopy, Sigma Aldrich, USA
* Centrifuge-•Universal 320 Hettich Zentrifugen, Tuttlingen, Germany
* Desiccator, 25 L

# Loading of quercetin into MMNPs

REAGENTS AND CHEMICALS REQUIRED

* Ethanol, 98%, TTT d.o.o. Sveta Nedjelja, Croatia
* Quercetin-•Lach-•ner, ≥ 99 %, Czechia APPARATUS REQUIRED
* Centrifuge-•Universal 320 Hettich Zentrifugen, Tuttlingen, Germany
* Desiccator, 25 L



* UV-•VIS Cary 50 Spectrophotometer, Agilent Technologies, USA
* ABB Bomem MB102 spectrometer, equipped with CsI optics and a DTGS detector
* Barnstead/Lab-•Line Model SHKE Max Q Mini 4450 E-•Class Benchtop High



* Temperature Incubator Shaker, Connecticut, USA
* Vortex V1 S000, IKA, Staufen, Germany
* Analytical balance, Phoenix Instrument, Garbsen, Germany
* Test tubes
* Eppendorf pipettes (10 µL, 100 µL, 1000 µL, 5 mL)
* Syringe Filter (F2613-•3, PTFE 0,45 µm. 17 mm), Macherey-•Nagel, Germany

# Release of flavonol with and without external magnetic fields

REAGENTS AND CHEMICALS REQUIRED

* Ethanol, 98%, TTT d.o.o. Sveta Nedjelja, Croatia
* Phosphate-•buffered saline tablet, Sigma Aldrich, USA
* Deionized water (Milipore Q water) APPARATUS REQUIRED
* Spectra/Por Closures for Standard RC Membrane MWCO 8kD 7,5/12mm, Spectra/Por® 6 Standard RC Pre-•wetted Dialysis Tubing 8 MWCO, 1 kD, Flat Widths 8 mm,
* Thermostat-•BTR-•2000-•A
* Centrifuge-•Universal 320 Hettich Zentrifugen, Tuttlingen, Germany
* Desiccator, 25 L
* UV-•VIS Cary 50 Spectrophotometer, Agilent Technologies, USA
* ABB Bomem MB102 spectrometer, equipped with CsI optics and a DTGS detector
* Analytical balance, Phoenix Instrument, Garbsen, Germany
* Scale
* Volumetric flask
* Test tubes
* Magnet
* Function generator, Iskra, Slovenia
* Wires
* Erlenmeyer flask, 10 mL, 25 mL, 50 mL, 100 mL
* Spatula
* Eppendorf pipettes (10 µL, 100 µL, 1000 µL 5 mL)
* Syringe Filter (F2613-•3, PTFE 0,45 µm. 17 mm), Macherey-•Nagel, Germany

# APPENDIX.2. Data accompanying graphs and conditions

|  |  |
| --- | --- |
| Frequency table | |
| Diameter/nm | Number of MMNPs |
| 27.800 -• 36.799 | 49 |
| 36.800 -• 45.799 | 96 |
| 45.800 -• 54.799 | 99 |
| 54.800 -• 63.799 | 105 |
| 63.800 -• 72.799 | 58 |
| 72.800 -• 81.799 | 41 |
| 81.800 -• 90.799 | 23 |
| 90.800 -• 99.799 | 17 |
| 99.800 -• 108.799 | 7 |
| 108.800 -• 117.799 | 5 |

**Table.2** . MMNPs diameter data obtained using Image-•J (Media Cybernetics, USA) Additional information involving histogram (Figure 6):

-• Standard deviation = 17.546

-• Mean = 51.620

-• Lowest Diameter = 27.873

-• Highest Diameter = 117.174

|  |  |  |  |
| --- | --- | --- | --- |
| Evacuation phase | | Heating phase | |
| The velocity of heating | 10,0 oC/min | Velocity | 10.0 oC/min |
| Temperature | 100 oC | Temperature | 110 oC |
| Velocity | 1333.22 Pa/s | Time | 720 min |
| Pressure | 0.0026664473  Pa |  |  |
| Time | 60 min |  |  |

**Table.3.** Experimental conditions under performed MMNPs

|  |  |  |  |  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
|  | f = 10 Hz | | | f = 0.1 Hz | | | f = 1 Hz | | | f = 0 Hz | | |
| t / min | c / mol dm-•3 ± 1 10-•11 | m / mg  ±  1 10-•7 | m / % | c / mol dm-•3 ± 1 10-•11 | m / mg  ± 1 10-•7 | m / % | c / mol dm-•3 ± 1 10-•11 | m / mg± 1 10-•7 | m / % | c / mol dm-•3  ± 1  10-•11 | m / mg± 1 10-•7 | m / % |
| 5 | 7.2826  10-•7 | 6.6031  10-•3 | 0.0929 | 39.3057  10-•7 | 35.6388  10-•3 | 0.1731 | 14.0509  10-•7 | 12.7400  10-•3 | 0.1051 | 18.091  7 10-•7 | 16.4038  10-•3 | 0.4544 |
| 10 | 26.6324  10-•7 | 24.477  9  10-•3 | 0.3443 | 21.0304  10-•7 | 20.8504  10-•3 | 0.1013 | 18.1835  10-•7 | 17.1241  10-•3 | 0.1413 | 19.561  0 10-•7 | 18.5563  10-•3 | 0.5140 |
| 20 | 19.1018  10-•7 | 18.857  3 10-•3 | 0.2652 | 27.4590  10-•7 | 27.6326  10-•3 | 0.1342 | 37.5608  10-•7 | 35.5181  10-•3 | 0.2931 | 37.928  2 10-•7 | 36.0968  10-•3 | 0.9999 |
| 30 | 79.9890  10-•7 | 74.930  2 10-•3 | 1.0539 | 36.4588  10-•7 | 37.0377  10-•3 | 0.1799 | 36.1466  10-•7 | 35.9386  10-•3 | 0.2965 | 39.489  4 10-•7 | 39.2318  10-•3 | 1.0868 |
| 40 | 175.222  7 10-•7 | 164.90  57 10-•3 | 2.3193 | 59.5096  10-•7 | 59.5909  10-•3 | 0.2894 | 60.8871  10-•7 | 60.0098  10-•3 | 0.4951 | 63.183  0 10-•7 | 62.5053  10-•3 | 1.7314 |
| 50 | 83.3869  10-•7 | 89.581  2 10-•3 | 1.2599 | 77.6012  10-•7 | 78.6927  10-•3 | 0.3822 | 78.2000  10-•7 | 70.9046  10-•3 | 0.5850 | 80.742  0 10-•7 | 81.2906  10-•3 | 2.2518 |
| 60 | 159.977  9 10-•7 | 162.80  73 10-•3 | 2.2898 | 67.8667  10-•7 | 73.3843  10-•3 | 0.3564 | 78.5196  10-•7 | 78.7576  10-•3 | 0.6498 | 98.723  5 10-•7 | 101.255  0 10-•3 | 2.8048 |

|  |  |  |  |  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
| 90 | 200.936  7 10-•7 | 207.19  76 10-•3 | 2.9142 | 118.651  9  10-•7 | 122.508  4 10-•3 | 0.5950 | 92.2031  10-•7 | 94.7243  10-•3 | 0.7816 | 123.15  18  10-•7 | 127.880  1 10-•3 | 3.5424 |
| 120 | 251.079  1  10-•7 | 261.77  17 10-•3 | 3.6817 | 131.876  2  10-•7 | 119.573  2 10-•3 | 0.5807 | 125.080  4 10-•7 | 113.411  4 10-•3 | 0.9357 | 138.30  47  10-•7 | 125.402  0 10-•3 | 3.4737 |
| 180 | 611.718  2  10-•7 | 600.14  89 10-•3 | 8.4409 | 213.977  4  10-•7 | 199.993  7 10-•3 | 0.9713 | 153.641  3 10-•7 | 144.978  3 10-•3 | 1.1962 | 153.09  03  10-•7 | 145.078  3 10-•3 | 4.0188 |
| 240 | 712.002  9  10-•7 | 645.57  88 10-•3 | 9.0799 | 251.455  6  10-•7 | 227.996  8 10-•3 | 1.1073 | 185.232  8 10-•7 | 180.588  0 10-•3 | 1.4900 | 182.47  77  10-•7 | 178.664  5 10-•3 | 4.9492 |

**Table 4**. Results for quercetin release

# APPENDIX.3. Extension

After doing the experiment involving quercetin for my Extended Essay, the project continued with a plan to carry out the above mentioned extension. Other flavonoids including myricetin and fisetin are being used as to see how they compare to quercetin. Since they are more expensive and considered to be more promising in future treatments in the world of healthcare, the experiments will be performed with them multiple times, and the experiment with quercetin will be repeated. As the results obtained showed that the hypothesis stands for 10 Hz, future experiments could and will probably be conducted by examining the release of flavonoids when even higher frequencies are applied.

Furthermore, an extension that was performed at the institute once I was writing my Extended Essay is the thermogravimetric analysis. Thermogravimetric analysis (TGA) is a technique which measures the change in mass of polymer as a function of temperature or time under certain and controlled conditions such as atmosphere, temperature and flow. The measuring procedure is usually done in an atmosphere such as N2, He, Ar, air or O2.The TGA analysis can be performed using an isothermal or a non-•isothermal method. The apparatus that records the mass change is called the thermo scale which has a bowl on which the sample is put and which then enters into a small electric oven that contains thermo element which measures the exact temperature. The instrument has to have high precision in measuring weight, temperature and time. Methods are divided into two groups: differential and integral. Non-•isothermal thermogravimetric analysis results in a thermogravimetric curve that represents mass change of the sample dependent on time and temperature. This is seen on Figure 14 below and is represented with a red curve. Deriving TGA curve, differential thermogravimetric curve (DTG) is obtained, which represents the speed of mass change of the sample with temperature or time. This is seen on Figure 14 below and is represented with

a blue curve. The shape of the curve depends on the conditions of the experiment: speed of heating, shape and mass of the sample, type of gas, which flows through the oven.41

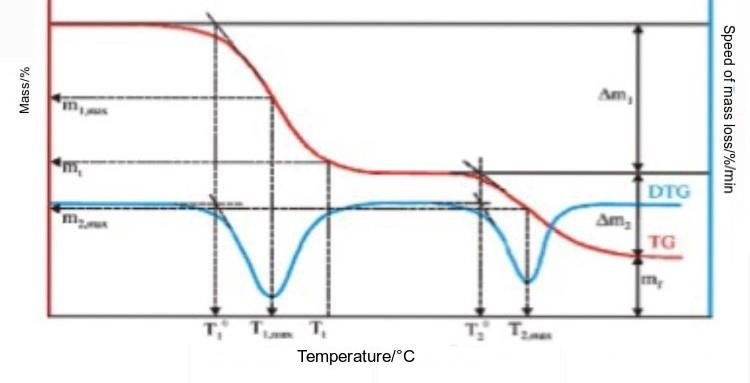


Figure.14:Thermogravimetric analysis

41 Notes written during the stay at the Ruder Boskovic Institute (Information given by Suzana Segota and other fellows)

**Appendix.4.**

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**RECOMMENDATION LETTER**

It is my pleasure to fill out a letter of support for Sara Dožai, student of the International Baccalaureate® (IB) Diploma Programme (DP) core in Zagreb. I met Sara Dožai in the Summer 2018, when Sara came to the Ruđer Bošković Institute for the purpose of her Extended Essay that is mandatory for all IB students. In the Summer of 2018, more specifically from 18th of June 2018 until 13th of July, Sara began to work at the Ruđer Bošković Institute in Laboratory of Biocolloids and Surface Chemistry. Sara began to work on nanostructure issues and introduced numerous spectroscopic methods for nanocharacterization of materials and surface processes within the project "Protective mechanisms and effects of nano-delivered flavonoids in model cell membranes and neurons (NanoFlavNeuroProtect)". She worked independently and followed directions when needed well. As her supervisor outside of school, I gave her all needed materials to study from and helped her with the research that needed to be done. The experiments in themselves have long methods which were designed with the help of all participants in our laboratory. During her visit, her great interest for surface chemistry as well as the practical experience in work with laboratory equipment enabled her further development and advancement in the field of chemistry, especially in physical chemistry and nanotechnology. She was one of the few students who were able to penetrate deep into the field of colloidal chemistry and techniques used in their research. Finally, Sara has an outstanding gift for team work. So far she has proven to be an extremely dedicated student with excellent understanding and clear vision of her goals. In this short time that she has spent working in our group she managed to do an extraordinary amount of work and obtain many new and interesting results.

It is therefore my great pleasure to recommend Sara Dožai as a good candidate for any scholarship, which will give her a valuable opportunity to show her work to the academic world.

